

Valences of Some Common Ions

Cations (positive charge)		Anions (negative charge)	
+1		-1	
H ⁺ hydrogen	H ₃ O ⁺ hydronium	H ⁻ hydride	MnO ₄ ⁻ permanganate
Li ⁺ lithium	Hg ⁺ mercury(I)	F ⁻ fluoride	ClO ⁻ hypochlorite
Na ⁺ sodium		Cl ⁻ chloride	ClO ₂ ⁻ chlorite
K ⁺ potassium		Br ⁻ bromide	ClO ₃ ⁻ chlorate
Ag ⁺ silver		I ⁻ iodide	ClO ₄ ⁻ perchlorate
Cu ⁺ copper(I)		OH ⁻ hydroxide	HCO ₃ ⁻ hydrogen carbonate
NH ₄ ⁺ ammonium		NO ₂ ⁻ nitrite	C ₂ H ₃ O ₂ ⁻ acetate
		NO ₃ ⁻ nitrate	HSO ₃ ⁻ hydrogen sulfite
		CN ⁻ cyanide	HSO ₄ ⁻ hydrogen sulfate
			H ₂ PO ₄ ⁻ dihydrogen phosphate
+2		-2	
Mg ²⁺ magnesium	Cr ²⁺ chromium(II)	O ²⁻ oxide	C ₂ O ₄ ²⁻ oxalate
Ca ²⁺ calcium	Co ²⁺ cobalt(II)	S ²⁻ sulfide	S ₂ O ₃ ²⁻ thiosulfate
Sr ²⁺ strontium	Mn ²⁺ manganese(II)	CO ₃ ²⁻ carbonate	HPO ₄ ²⁻ monohydrogen phosphate
Ba ²⁺ barium	Cd ²⁺ cadmium	SO ₃ ²⁻ sulfite	
Zn ²⁺ zinc	Sn ²⁺ tin(II)	SO ₄ ²⁻ sulfate	O ₂ ²⁻ peroxide
Cu ²⁺ copper(II)	Pb ²⁺ lead(II)	CrO ₄ ²⁻ chromate	
Fe ²⁺ iron(II)	Ni ²⁺ nickel(II)	Cr ₂ O ₇ ²⁻ dichromate	
Hg ²⁺ mercury(II)			
+3		-3	
Al ³⁺ aluminum	Cr ³⁺ chromium(III)	N ³⁻ nitride	PO ₄ ³⁻ phosphate
Fe ³⁺ iron(III)	Co ³⁺ cobalt(III)	P ³⁻ phosphide	
+4			
Sn ⁴⁺ tin(IV)			
Pb ⁴⁺ lead(IV)			