COMMON IONS

Polyatomic ions are **bolded.**

CATIONS (positive)

1+

H^+	hydrogen			**H_3O^+**	**hydronium**
Li^+	lithium	Cu^+	copper (I)	**NH_4^+**	**ammonium**
Na^+	sodium		*cuprous*	**Hg_2^{2+}**	**mercury (I)**
K^+	potassium	Au^+	gold (I)		***mercurous***
Rb^+	rubidium		*aurous*		
Cs^+	cesium				
Ag^+	silver				

2+

Be^{2+}	beryllium	Cr^{2+}	chromium (II)	Pb^{2+}	lead (II)
Mg^{2+}	magnesium		*chromous*		*plumbous*
Ca^{2+}	calcium	Co^{2+}	cobalt (II)	Mn^{2+}	manganese (II)
Sr^{2+}	strontium		*cobaltous*		*manganous*
Ba^{2+}	barium	Cu^{2+}	copper (II)	Hg^{2+}	mercury (II)
Ra^{2+}	radium		*cupric*		*mercuric*
Cd^{2+}	cadmium	Fe^{2+}	iron (II)	Ni^{2+}	nickel*
Ni^{2+}	nickel*		*ferrous*	Pt^{2+}	platinum (II)
Zn^{2+}	zinc			Sn^{2+}	tin (II)
					stannous

3+

Al^{3+}	aluminum	Cr^{3+}	chromium (III)	Fe^{3+}	iron (III)
Sb^{3+}	antimony		*chromic*		*ferric*
Bi^{3+}	bismuth	Co^{3+}	cobalt (III)	Mn^{3+}	manganese (III)
Ga^{3+}	gallium		*cobaltic*		*manganic*
		Au^{3+}	gold (III)	Ti^{3+}	titanium (III)
			auric	U^{3+}	uranium (III)

4+

Pb^{4+}	lead (IV)	Sn^{4+}	tin (IV)	W^{4+}	tungsten (IV)
	plumbic		*stannic*	U^{4+}	uranium (IV)
Pt^{4+}	platinum (IV)	Ti^{4+}	titanium (IV)	V^{4+}	vanadium (IV)

ANIONS (negative)

1–

H^-	hydride	**$C_2H_3O_2^-$** (or **CH_3COO^-**)	**acetate**	**CNO^-**	**fulminate**
F^-	fluoride	**NH_2^-**	**amide**	**OH^-**	**hydroxide**
Cl^-	chloride	**$C_6H_5COO^-$**	**benzoate**	**IO^-**	**hypoiodate**
Br^-	bromide	**BrO^-**	**hypobromite**	**IO_3^-**	**iodate**
I^-	iodide	**BrO_2^-**	**bromite**	**IO_4^-**	**periodate**
		BrO_3^-	**bromate**	**MnO_4^-**	**permanganate**
		BrO_4^-	**perbromate**	**NO_2^-**	**nitrite**
		CN^-	**cyanide**	**NO_3^-**	**nitrate**
		OCN^-	**cyanate**	**$HC_2O_4^-$**	**binoxalate**
		HCO_3^-	**bicarbonate** (or **hydrogen carbonate**)	**SCN^-**	**thiocyanate**
				$H_2PO_4^-$	**dihydrogen phosphate**
				HS^-	**bisulfide** (or **hydrogen sulfide**)
		ClO^-	**hypochlorite**	**HSO_3^-**	**bisulfite** (or **hydrogen sulfite**)
		ClO_2^-	**chlorite**	**HSO_4^-**	**bisulfate** (or **hydrogen sulfate**)
		ClO_3^-	**chlorate**		
		ClO_4^-	**perchlorate**		

2–

O^{2-}	oxide	**$B_4O_7^{2-}$**	**tetraborate**	**$C_2O_4^{2-}$**	**oxalate**
O_2^{2-}	**peroxide**	**CO_3^{2-}**	**carbonate**	**HPO_4^{2-}**	**hydrogen phosphate**
S^{2-}	sulfide	**$C_4H_4O_6^{2-}$**	**tartrate**	**$S_2O_3^{2-}$**	**thiosulfate**
Se^{2-}	selenide	**CrO_4^{2-}**	**chromate**	**SeO_4^{2-}**	**selenate**
Te^{2-}	telluride	**$Cr_2O_7^{2-}$**	**dichromate**	**SiO_4^{2-}**	**silicate**
		SO_2^{2-}	**hyposulfite**	**SiF_6^{2-}**	**hexafluorosilicate**
		SO_3^{2-}	**sulfite**		
		SO_4^{2-}	**sulfate**		

3–

N^{3-}	nitride	**AsO_4^{3-}**	**arsonate**	**PO_3^{3-}**	**phosphite**
P^{3-}	phosphide	**BO_3^{3-}**	**borate**	**PO_4^{3-}**	**phosphate**
As^{3-}	arsenide	**$C_2H_5O_7^{3-}$**	**citrate**		

The most common oxidation state of nickel is +2, but compounds of Ni^+, Ni^{3+}, and Ni^{4+} are known.

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